

The ESR spectra of aqueous solutions of the complexes were recorded at 300 K on a SPINSCAN X radio spectrometer (ADANI, Belarus), in TABLE 1 parameters are mention, while in FIG. 1-3 parameters are shown.

TABLE 1. ESR parameters of the spectra of complexes I-III.

№	Complex compounds	A, €	g
1	(Cu ₂ Gln ₂ ·2Urea)·2L (I)	55,44	2,1212
2	(Cu ₂ Gln ₂ ·2Atsetamid)·2L (II)	56,24	2,1210
3	(Cu ₂ Gln ₂ ·2Tiourea)·2L (III)	57,35	2,1208
4	(CuGln ₂)·2CH ₃ COO (IV)	90	2,0495

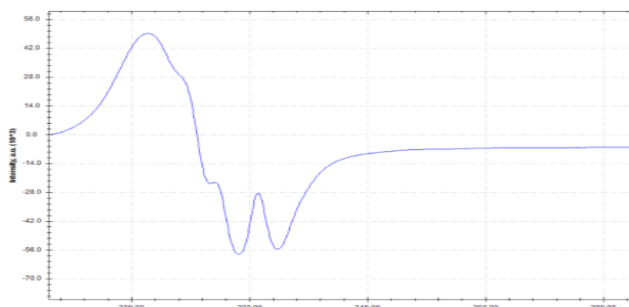


FIG. 1. ESR spectrum of the complex I

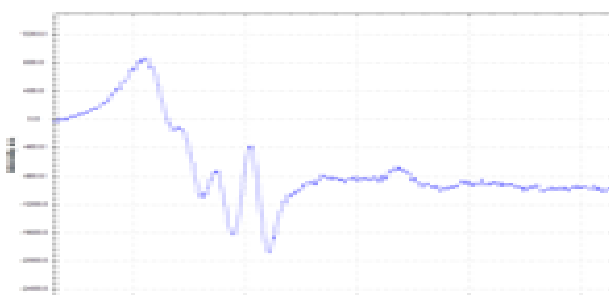


FIG. 2. ESR spectrum of the complex II

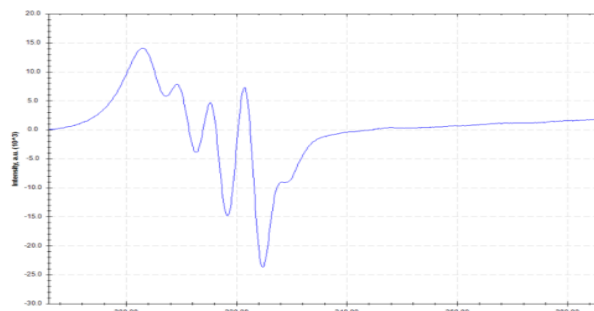


FIG. 3. ESR spectrum of the complex III

When studying the intramolecular complex of glutamine with copper(II) ions, the g-factor was determined to be 2.062 and 2.037 (TABLES 1 and 2) (FIG. 4). Some papers have predicted that the g-factor varies with donor groups and their coordination. TABLE 2 shows these g-factors of different glutamine complexes.

TABLE 2. ESR parameters of the spectra of copper complexes IV [1-3].

Literature	In thus work	[1] 9.44 GHz	[2] 9.8 GHz	[2] 33.9 GHz	[3] 33.9 GHz	[5]
g-factor of the copper(II) glutamine complex	2,062 2,037	2,063 2,274	2,049 2,248	2,053 2,248	2,254 2,295	2,039 2,217

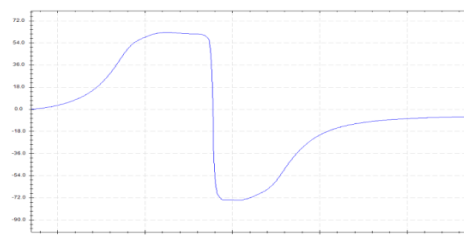


FIG. 4. ESR spectrum of the complex IV.

It is concluded that in Na, Mg salts glutamine is in the carboxylate form with the formation of an ionic bond with the α -COO⁻ group and molecular complexes, the ions Zn²⁺, Cu²⁺, Ni²⁺ are coordinated by the bidentate-cyclic function of the α -COO⁻ group. In the intra-molecular complex of the copper (II) ion, the glutamine ligand is bidentately coordinated through the N and O atoms of the α -carboxylate group to form a five-membered chelate cycle around the central ion.

Based on the results of studying the IR spectra, a conclusion was made about the internal molecular complex of the glutamine complex of the copper(II) ion (FIG. 5), in which the amino acid anion is connected bidentately through the N and O atoms of the α -hydrocarboxylate group to form a chelate ring around the Cu²⁺ ion [4 -6].

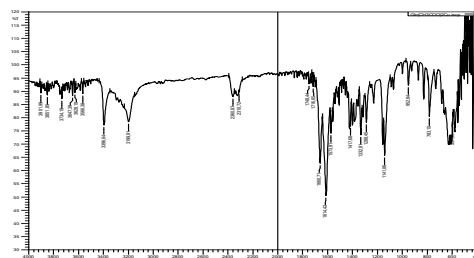


FIG 5. IR spectrum of the complex IV

Based on ESR spectroscopy data, the formation of monomeric (IV) and dimeric complexes (I-III) of copper (II) has been established. It is noted that the formation of dimeric complexes is associated with the participation of L1-3 amide ligands and a monomer complex in the presence of a second carboxylate ligand. Identification of the complexes was carried out based on diffractograms taken on the XRD-6100 diffractometer (FIG. 6 and 7).

In order to identify practical applications, the biological activity of synthesized complexes of the composition [M⁺ⁿ(Gln)_xL⁴] (IV-VIII) concerning two different pathogenic bacteria *Pseudomonas aeruginosa* and *Candida albicans*.

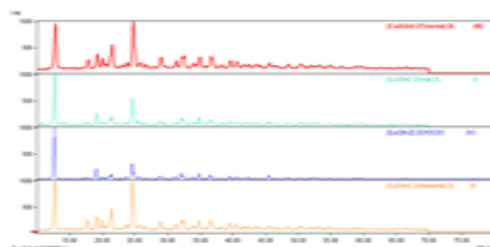


FIG. 6. Compared XRD spectrum of the complexes I-IV.

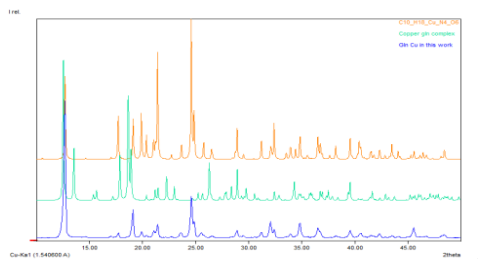


FIG. 7. Compared XRD spectrum of the complexes IV.

Microorganisms used to study the antibacterial activity of 3d metal glutamine complexes are *Pseudomonas aeruginosa* and *Candida albicans*. *Pseudomonas aeruginosa* causes a wide range of infections in humans, such as urinary tract infections, respiratory system infections, and soft tissue infections. *Candida albicans* do not cause problems at normal levels, and when it grows out of control, the level of healthy bacteria is minimized; it causes fungal infections in humans. **FIG. 8** and **9** shows the inhibitory activity of two different pathogenic bacteria, *Pseudomonas aeruginosa* and *Candida albicans*. It was found that some complexes have a higher antagonism to the control, for example, the copper(II) ion complex in both bacteria of the same equal growth of 24 mm-26 mm. And the complex of the cobalt(III) ion is more antagonistically equal to the growth of 36 mm-38 mm/26 mm-28 mm, for chromium(III) the least than the complex of copper(II) is equal to the growth of 13 mm-15 mm/19 mm-20 mm.

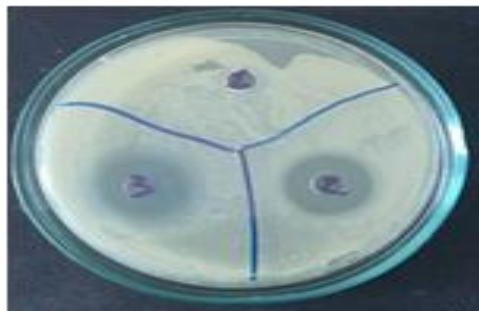


FIG. 8. Biological activity with *P. aeruginosa*.

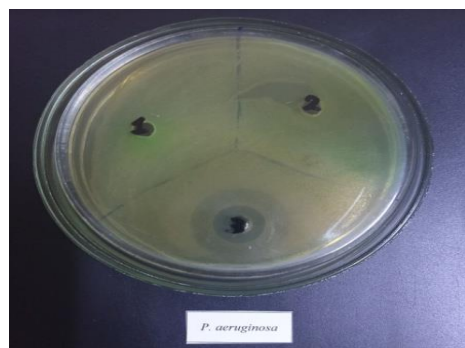


FIG. 9. Biological activity with *C. albicans*.

It was revealed that free glutamine and complexes of Mn^{+2} , Ni^{+2} ions did not exhibit antibacterial properties to both types of bacteria. Complexes of Cu^{2+} , Co^{2+} , Cr^{3+} ions showed antibacterial activity, and copper complexes showed almost the same antagonism against both bacteria, cobalt complexes against *Pseudomonas aeruginosa* were 1.37 times more active than *Candida albicans*. Chrome complexes against the bacterium *Candida albicans* showed activity 1.39 times more than against the bacterium *Pseudomonas aeruginosa*. In general, according to the antibacterial activity, the complexes, depending on the nature of the complexing agents, were located in the series $Co^{2+} > Cu^{2+} > Cr^{3+}$.

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