

CuO Nanoparticles Synthesis by Sol- gel Method and Characterization

Madhav Jadhav *

Department of Physics, SMBS Thorat College, Maharashtra, India

*Corresponding author: Madhav Jadhav, Department of Physics, SMBS Thorat College, Maharashtra, India, Tel: +9881800832; E-mail: mjsmbst2008@gmail.com

Received: September 23, 2021; Accepted: October 07, 2021; Published: October 14, 2021

Abstract

CuO one of the useful metal oxide and has many applications in different fields. CuO acts as a semiconductor. In the present work, the main objective is to synthesize CuO nanoparticles using a low cost Sol- Gel process at room temperature and study the material by using the characterization techniques XRD and UV-visible spectrum. Crystallites size was calculated from Scherer's formula using most peak of CuO. The XRD pattern of CuO shows the highest intensity diffraction peak at 38.540 correspond to (111) plane. The average crystallite sizes of CuO particles were found to be 16.57 nm. CuO nanoparticle showed maximum absorption at wavelength 322 nm. The calculated energy band gap of CuO were 3.14 eV.

Keywords: CuO nanoparticles; Sol-Gel method

Introduction

Nanoparticles are different from bulk materials and isolated molecules because of their unique optical and, electronic and chemical properties. In the universe there is number of metal oxides are available in nature but some of the metal oxides are most useful in accordance with their applications in science and technology [1]. Some metal oxides like TiO₂, ZnO and Fe₃O₄ are proved for number of applications. Similarly, CuO is also one of the useful metal oxide and has many applications in different fields.

CuO acts as a semiconductor and has energy band gap nearly 2.6 eV. Semiconductor materials are interesting because of their great practical importance in the devices such as electrochemical cell, gas sensors, magnetic storage devices, field emitters, high T_c -superconductors, Nano fluids, and catalysts etc. in the past decades, various methods have been proposed to produce CuO nanoparticles with different sizes and shapes such as thermal oxidation, sonochemical, combustion and quick precipitation [2]. The properties of nanomaterials will depend mostly on the Nano powders size, morphology and specific surface area of the prepared materials. Some aspects are strongly depending on the preparation methods. In the present work, the main objective is to synthesize CuO nanoparticles using a low cost Sol-Gel process at room temperature and study the material by using the characterization techniques XRD and UV-visible spectrum.

Materials and Methods

Experimental details

The preparation was carried out in room temperature. All chemicals were used analytical grade without further purification. Double distilled water was used for all the experiments. CuCl_2 , NaOH, Acetic acid chemicals were used. CuO Nano particles were prepared by Sol-gel method. The aqueous solution of $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ (0.2M) (M.W. 170.48) is prepared in cleaned Beaker. 1 ml of glacial acetic acid is added to above aqueous solution and heated to 80°C with constant stirring. 8 M. NaOH is added to above heated solution till PH reaches to 10.2. The colour of solution turned from blue to black immediately and the large amount of black precipitate is formed immediately. The obtained precipitate was dried in air for 24 hrs. This powder is further used for the characterization of CuO nanoparticles. The crystal phase of CuO nanoparticles were determined by X-Ray Diffractometer (XRD) and band gap was determined by UV-Visible spectrum.

Results and Discussion

The structural analysis of CuO nanopowder was recorded using an Burker D-8 X-Ray diffractometer operated at 30 KV and 20 mA. The XRD pattern of CuO nanopowder. The CuO diffraction pattern shows broad peaks because CuO with small sizes were formed. The peak observed at $2\theta = 35.540$, 38.540 , and 48.750 . Corresponding to the plane (002), (111), and (202) respectively [3]. Crystallites size was calculated from Scherer's formula using most peak of CuO. the XRD pattern of CuO shows the highest intensity diffraction peak at 38.540 correspond to (111) plane. The average crystallite sizes of CuO particles were found to be 16.57 nm [4]. It is observed that the CuO concentration increases in composites the intensity of CuO. The observed diffraction peak patterns were compared with standard values of reference No. 00-045-0937 (Figures 1-3).

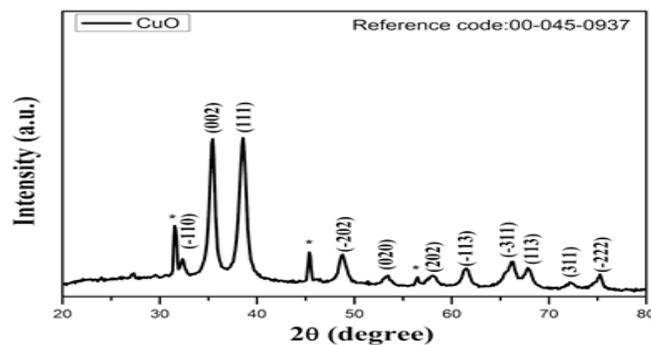


FIG.1. XRD Pattern of CUO.

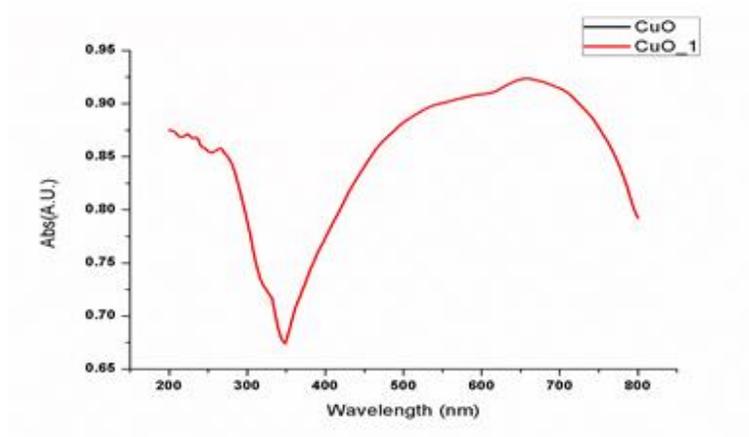


FIG.2. UV graph of CUO.

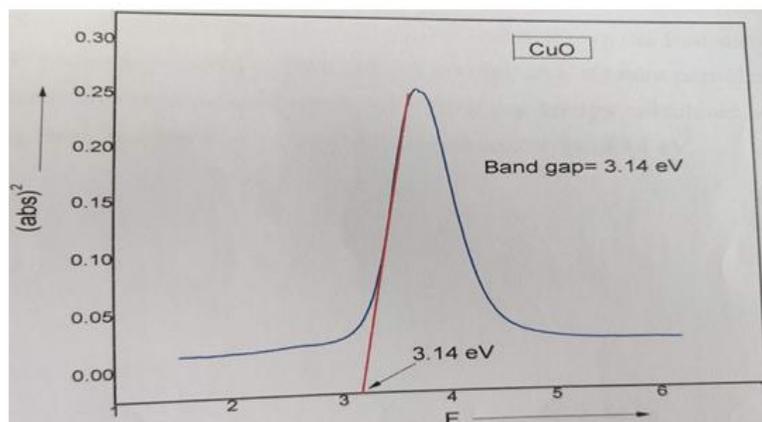


FIG.3. Uv -Visible spectra of CuO 1.

Optical absorption Studies

The Uv-Visible absorption spectra of nano powder CuO are recorded at room temperature are given. The UV- Visible spectra of CuO nanoparticles measured in the range 200 nm-800 nm [5]. The band gap energy calculated from $E=1240/\lambda$ relation by plotting the absorption edge of CuO were found to be 3.14 eV. CuO nanoparticles showed maximum absorption at wavelength 322 nm (Table 1).

TABLE.1. Peak List.

Pos. [$^{\circ}2\theta$.]	Height [cts]	FWHM [$^{\circ}2\theta$.]	d-spacing [Å]	Rel. Int. [%]
27.2249	370.93	0.2952	3.27566	2.63
31.5842	5213.74	0.2952	2.83279	36.94
32.3319	1689.32	0.2952	2.76897	11.97
35.4238	14052.69	0.492	2.53405	99.56
38.5603	14114.65	0.6888	2.33484	100
45.3814	3077.9	0.2952	1.99851	21.81
48.7349	2888.5	0.5904	1.86854	20.46
53.3706	861.68	0.6888	1.71667	6.1
56.4356	799.75	0.2952	1.63049	5.67
58.1988	1014.83	0.5904	1.58522	7.19
61.4851	1897.29	0.6888	1.50815	13.44
66.2288	2596.68	0.492	1.41116	18.4
67.8662	1953.2	0.5904	1.38105	13.84
72.2221	586.14	0.7872	1.30811	4.15
75.2472	1422.92	0.36	1.26181	10.08

Conclusion

Strain of the prepared nanoparticles calculation concluded that the small particle has high strain and big particles have low strain. The average particle size determined from particle analyzer is given by 16.57 nm. The UV-visible pattern of CuO nanoparticles showed maximum absorption at 322 nm. The absorption wavelength of CuO nano powder shows the red shift region and there is enhancement in the band gap value of CuO nanoparticles when compared with bulk copper oxide. The band gap energy calculated by extra plotting the absorption edge of CuO was found to be 3.14 eV.

REFERENCES

1. Jayaprakash J, Srinivasan N, Chandrasekaran P, Girija EK. Synthesis and characterization of cluster of grapes like pure and Zinc-doped CuO nanoparticles by sol-gel method. *Spectrochimica Acta A Mol Biomol Spectrosc.* 2015;136:1803-06.
2. Kayani ZN, Umer M, Riaz S, Naseem S. Characterization of copper oxide nanoparticles fabricated by the sol-gel method. *J Elect Materi.* 2015;44(10):3704-9.
3. Sahoo M, Sabbaghi S, Saboori R. Synthesis and characterization of mono sized CuO nanoparticles. *Materi Letters.* 2012;81:169-72.
4. Nithya K, Yuvasree P, Neelakandeswari N, Rajasekaran N, et al. Preparation and characterization of copper oxide nanoparticles. *Int J Chem Tech Res.* 2014;6(3):2220-2.
5. Lanje AS, Sharma SJ, Pode RB, Ningthoujam RS. Synthesis and optical characterization of copper oxide nanoparticles. *Adv Appl Sci Res.* 2010;1(2):36-40.