



## **STUDIES ON SYNTHESIS OF 4 – OXOQUINAZOLIN DYES AND THEIR APPLICATION ON VARIOUS FIBRES**

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### **ABSTRACT**

A series of 3-[4-(4-aryalazo)-phenyl]-2-phenyl-3H-quinazolin-4-one derivatives (**3**) have been obtained by a reaction of a various coupling agents with diazonium salt containing 4-oxoquinazolin moiety (**2**). The diazonium salt (**2**) is obtained by the reaction of 4-(2-phenyl-4-oxo-3-quinazoliny)-aniline with NaNO<sub>2</sub> and HCl. The product is characterized by spectral and analytical data. Most of the tested compounds show promising dyeing properties.

**Key words :** 4-Oxoquinazolin, Diazotization, Dyeing.

### **INTRODUCTION**

The wide variety of 4-oxoquinazolin derivatives are now a major group of dyes and have soon attained a commercial status. There is no slackening of activity in this field as seen from the large number of patent specification and several ranges, which continue to appear in the market<sup>1-3</sup>. The use of the dyestuff makes possible the highest degree of fastness to severe washing, abrasion, etc. At the same time, the shade ranges that can be achieved on cotton with fast dyestuff has considerably been extended<sup>4</sup>. Improvements in the structure of reactive dye chromogens and in the structure selection and number of reactive group have led to an increased use of reactive dyes<sup>5-8</sup>. Patel et al.<sup>9</sup> have synthesized fiber reactive dyes for silk, wool and rayon.

We report here the synthesis and study of the dyeing properties of the 4-oxoquinazolin dyes based on 2-phenyl-3, 1-benzoxazine-4(4H)-one.

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The reaction of benzoyl chloride with anthranilic acid in pyridine at 8°C gave 2-phenyl-3,1-benzoxazine-4(4H)-one (**1**). The compound (**1**) on condensation with 1,4-diaminobenzene yielded 4-(2-phenyl-4-oxo-3-quinazolinyl)-aniline (**2**). Compound (**2**) diazotized and coupled with different coupler (**a-o**) gives different types of dyes (**3**). All the compounds synthesized were adequately characterized by their elemental analysis and spectral data.

## EXPERIMENTAL

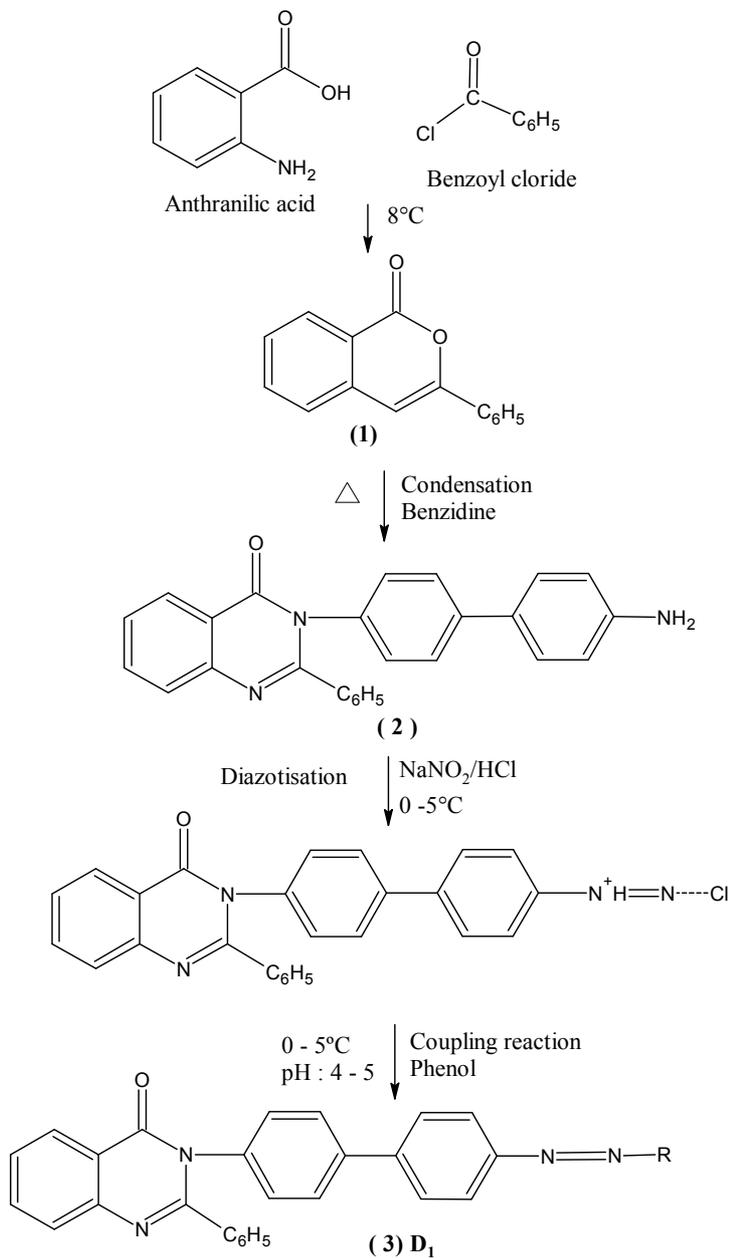
Melting points were taken in open capillaries and are uncorrected. The IR spectra of dyes D<sub>1</sub> to D<sub>15</sub> were recorded on Bio-Red FTS-40 spectrophotometer using KBr pellets. The purity of all dyes has been checked by thin-layer chromatography<sup>10</sup>. The absorption spectra of all the dyes were recorded on Beckmann DB-GT Grafting Spectrophotometer. Fastness to light was assessed in accordance with Bs : 1006-1978. The rubbing fastness was carried out with a crock meter (Atlas) in accordance with AATCC (1961) and the wash fastness test in accordance with IS : 765-1979.

### 2-Phenyl-3,1-benzoxazine 4(4H)-one<sup>10</sup> (**1**)

Benzoyl chloride (140.5 g; 1M) was added dropwise to anthranilic acid (137.0 g; 1M) dissolved in pyridine (60 mL) with constant stirring at 8°C over the period of 1 hour. After the addition of benzoyl chloride, reaction mixture was stirred for half an hour at room temperature. At the end of the reaction, reaction mixture almost solidified. The solid mass was poured into cold water, filtered, washed successively with aqueous sodium bicarbonate solution (10% to remove unreacted anthranilic acid) and water, dried and recrystallised from ethanol (95%) to get compound (**1**). Yield 84%, m. p. 118°C. Anal. Calcd. for C<sub>14</sub>H<sub>9</sub>O<sub>2</sub>N : C, 75.33; H, 4.03; N, 6.27. Found C, 75.35; H, 4.00; N, 6.25%.

### 4-(2-Phenyl-4-oxo-3-quinazolinyl)-aminodiphenyl<sup>11</sup> (**2**)

Equimolar ratio of compound (**1**) (223.0 g; 1M) and 1,4-diaminobenzene (108.0 g; 1M) were intimately mixed and heated on a free flame for five minutes with vigorous shaking. To the hot reaction mixture ethanol (750.0 mL) was added and the contents of the flask were allowed to cool. Scratching the side with a glass rod yielded a yellowish-white crystalline solid. It was filtered, washed with cold ethanol and recrystallised from ethanol (95%) to get compound (**2**). Yield 72%, m. p. 210°C. Anal. Calcd. for C<sub>20</sub>H<sub>15</sub>ON<sub>3</sub> : C, 76.67; H, 4.79; N, 13.41. Found C, 76.65; H, 4.75; N, 13.40%.



Where R = Phenol, o-Cresol, m-Cresol, p-Cresol, o-Cl-Phenol, m-Cl-Phenol, l-Naphthol, 2-Naphthol, Salicylic acid, Resorcinol, H-acid, J-acid, R-acid and Gama acid for D<sub>1</sub> to D<sub>15</sub>, respectively

**Scheme**

**3 – [ 4 - (4 – Hydroxy - phenylazo) – phenyl ] – 2 – phenyl - 3H – quinazolin – 4 - one (D<sub>1</sub> to D<sub>15</sub>) : (3)**

Equimolar ratio of compound **(2)** (0.1945 g; 0.05 M) was suspended in water (10.0 mL), conc. hydrochloric acid (5.0 mL; 0.025M) was added drop-wise to the well stirred suspension and the solution was cooled to 0-5°C in an ice bath. A solution of sodium nitrite (5.0 mL; 10% W/V) was then added and the reaction mixture was stirred until a positive test for nitrous acid on starch-iodide paper (i. e., blue color on SI paper). The excess nitrous acid was neutralized with urea (1.0 g) and the mixture was filtered to get a clear diazonium salt solution which was used for the subsequent coupling reaction.

Phenol (0.188 g; 0.05 M) was dissolved in sodium hydroxide solution (15.0 mL; 5% W/V) and the solution was cooled to 0-5°C in an ice-bath. To this well-stirred solution, the above mentioned diazo solution was then gradually added in 1 hr. at 0-5°C maintaining pH 4-5 by the addition of the concentrated hydrochloric acid slowly and with vigorous stirring to the cold mixture until it is strongly acidic to litmus paper. The mixture was stirred for 3-4 hrs. at 0-5°C until all the diazo salt was consumed (spot test with alkaline phenol solution). After being stirred for further 2 hrs. to complete the separation, the dye was isolated by filtration, washed with ice water, dried and crystallized from ethanol (95%) to get orange crystals of compound **(3D<sub>1</sub>)**. Yield 77%, m. p. 169°C. Anal. Calcd. for C<sub>26</sub>H<sub>18</sub>O<sub>2</sub>N<sub>4</sub>: C, 74.64; H, 4.13; N, 13.40. Found C, 74.65; H, 4.10; N, 13.39%. IR : 1669 cm<sup>-1</sup> due to >C = O and at 1625 cm<sup>-1</sup> due to >C = N -. The absorption at 3429 cm<sup>-1</sup> is due to N - H and at 3247 cm<sup>-1</sup> is due to O - H. The absorption at 706 and 760 cm<sup>-1</sup> is due to mono substituted and at 846 cm<sup>-1</sup> is due to 1, 4-disubstituted benzene ring. The absorption at 1592 cm<sup>-1</sup> is due to - N = N -.

Other compounds **(3D<sub>2</sub>-D<sub>15</sub>)** were synthesized similarly from **(3)**, respectively. Characterization data are presented in Table 1.

## RESULTS AND DISCUSSION

All the dyes D<sub>1</sub> to D<sub>15</sub> were applied on nylon and polyester fibers using the reported procedure<sup>11-14</sup>. All the dyes were black, brown, violet, yellow to orange and obtained in excellent yields. Data on λ<sub>max</sub> value (in DMF solvent) and the results of exhaustion and fixation of all the dyes on nylon and polyester fabrics are furnished in Table 2.

Table 1. Characterization

Dye No.	Shade on dyed fibres	R	Molecular formula	Yield (%)	M. P. (°C)	Found (%) (Calcd.)		
						C	N	S
D <sub>1</sub>	Brown	(a) Phenol	C <sub>32</sub> H <sub>22</sub> O <sub>2</sub> N <sub>4</sub>	73	105	77.69 (77.73)	11.3 (11.34)	- (-)
D <sub>2</sub>	Green	(b) o-Cresol	C <sub>33</sub> H <sub>24</sub> O <sub>2</sub> N <sub>4</sub>	78	197	77.89 (77.95)	11.1 (11.03)	- (-)
D <sub>3</sub>	Brown	(c) m-Cresol	C <sub>32</sub> H <sub>24</sub> O <sub>2</sub> N <sub>4</sub>	74	145	77.9 (77.95)	11.01 (11.03)	- (-)
D <sub>4</sub>	Orange	(d) p-Cresol	C <sub>32</sub> H <sub>24</sub> O <sub>2</sub> N <sub>4</sub>	78	235	77.91 (77.95)	11.02 (11.03)	- (-)
D <sub>5</sub>	Orange	(e) o-Cl-Phenol	C <sub>32</sub> H <sub>21</sub> O <sub>2</sub> N <sub>4</sub> Cl	74	>300	72.65 (72.66)	10.5 (11.03)	- (-)
D <sub>6</sub>	Orange	(f) m-Cl-Phenol	C <sub>32</sub> H <sub>21</sub> O <sub>2</sub> N <sub>4</sub> Cl	78	>300	72.69 (72.66)	10.52 (10.60)	- (-)
D <sub>7</sub>	Orange	(g) p-Cl-Phenol	C <sub>32</sub> H <sub>21</sub> O <sub>2</sub> N <sub>4</sub> Cl	77	>300	72.63 (72.66)	10.51 (10.60)	- (-)
D <sub>8</sub>	Violet-black	(h) 1-Naphthol	C <sub>36</sub> H <sub>24</sub> O <sub>2</sub> N <sub>4</sub>	81	>300	79.4 (79.41)	10.31 (10.29)	- (-)
D <sub>9</sub>	Green	(i) 2-Naphthol	C <sub>36</sub> H <sub>24</sub> O <sub>2</sub> N <sub>4</sub>	84	>300	79.42 (79.41)	10.28 (10.29)	- (-)
D <sub>10</sub>	Orange	(j) Salicylic acid	C <sub>33</sub> H <sub>22</sub> O <sub>3</sub> N <sub>4</sub>	74	>300	75.81 (75.86)	10.7 (10.73)	- (-)
D <sub>11</sub>	Brown	(k) Resorcinol	C <sub>32</sub> H <sub>22</sub> O <sub>3</sub> N <sub>4</sub>	85	240	75.19 (75.29)	10.81 (10.98)	- (-)
D <sub>12</sub>	Green	(l) H-Acid	C <sub>36</sub> H <sub>23</sub> O <sub>8</sub> N <sub>5</sub> S <sub>2</sub> Na <sub>2</sub>	83	>300	56.58 (56.62)	9.15 (9.17)	8.35 (8.39)
D <sub>13</sub>	Violet	(m) J-Acid	C <sub>36</sub> H <sub>24</sub> O <sub>5</sub> N <sub>5</sub> SNa	87	>300	65.29 (65.62)	10.6 (10.59)	4.79 (4.84)
D <sub>14</sub>	Violet-Black	(n) R-Acid	C <sub>36</sub> H <sub>22</sub> O <sub>8</sub> N <sub>4</sub> S <sub>2</sub> Na <sub>2</sub>	84	>300	57.74 (57.75)	7.49 (7.48)	8.51 (8.55)
D <sub>15</sub>	Violet	(o) Gama-Acid	C <sub>36</sub> H <sub>24</sub> O <sub>3</sub> N <sub>5</sub> SNa	88	>300	65.4 (65.36)	10.55 (10.59)	4.81 (4.84)

**Table 2. Evaluation of exhaustion and fixation study of dyes on nylon and polyester fibers (N = Nylon, P = Polyester)**

Dye No.	$\lambda_{\max}$ (nm)	Log $\Sigma$	% Exhaustion		$\lambda^*_{\max}$ (nm)	% Fixation**	
			N	P		P	N
D <sub>1</sub>	446	4.15	53	51	454	73	70
D <sub>2</sub>	478	4.30	65	60	479	74	70
D <sub>3</sub>	443	4.20	66	62	448	70	68
D <sub>4</sub>	458	4.27	60	58	460	79	68
D <sub>5</sub>	720	4.98	62	58	721	69	65
D <sub>6</sub>	660	4.80	68	65	665	68	66
D <sub>7</sub>	600	4.91	59	55	609	66	60
D <sub>8</sub>	472	4.40	65	63	475	79	72
D <sub>9</sub>	480	4.38	60	54	482	82	77
D <sub>10</sub>	469	4.40	65	63	475	79	72
D <sub>11</sub>	480	4.38	60	55	482	82	77
D <sub>12</sub>	545	4.50	59	62	548	64	60
D <sub>13</sub>	560	4.36	57	52	562	75	70
D <sub>14</sub>	545	4.49	60	64	549	65	62
D <sub>15</sub>	536	4.70	70	66	539	75	66

The data reveals that the percentage exhaustion on nylon fibers is higher which may be due to the relatively open structure of the nylon fiber. The results of fastness to light, washing, rubbing, perspiration and sublimation of nylon and Polyester fibers are shown in Table 3. The light fastness of all the dyes on both the fibers was found to be fair to fairly good to good.

Table 3. Evaluation of fastness properties of dyed nylon and polyester patterns with dyes

Dye No	Light fastness		Wash fastness		Rubbing fastness				Perspiration fastness				Sublimation fastness	
	N	P	N	P	Dry		Wet		Acidic		Alkaline		N	P
					N	P	N	P	N	P	N	P		
D <sub>1</sub>	3	3	4	5	4	5	4	5	5	5	4	5	4	5
D <sub>2</sub>	3	3-4	5	5	4	5	4	4	4	5	4	5	4	5
D <sub>3</sub>	4	4	5	5	5	5	5	5	5	5	5	5	5	5
D <sub>4</sub>	3	3	5	5	4	4	5	4	4	4	4	4	4	4
D <sub>5</sub>	4-5	4-5	5	5	5	5	4	5	5	5	5	5	5	5
D <sub>6</sub>	3-4	3-4	5	5	4	5	5	5	4	5	4	5	5	5
D <sub>7</sub>	3-4	3-4	5	5	5	5	4	5	5	5	5	5	5	5
D <sub>8</sub>	4	4	5	5	5	5	4	5	4	5	4	5	4	5
D <sub>9</sub>	4	4-5	5	5	5	5	5	5	5	5	5	5	5	5
D <sub>10</sub>	4	4	5	5	4	5	4	5	4	5	4	5	5	5
D <sub>11</sub>	4	4	4	4	5	5	5	5	5	5	5	5	5	5
D <sub>12</sub>	4-5	4-5	5	5	5	5	5	5	5	5	5	5	5	5
D <sub>13</sub>	3	3-4	4	4	4	4	4	5	4	4	4	4	4	4
D <sub>14</sub>	4	4	5	5	4	5	4	5	5	5	5	5	5	5
D <sub>15</sub>	4	4	5	5	4	5	4	4	5	5	5	5	5	5

The obtained results of washing fastness of the prepared dyes for both the fibers showed that they are very good to excellent. Fastness to rubbing of dyed patterns was very good to excellent for all the dyes on both the fibers. This is attributed to good penetration and affinity of present dyes to synthetic fibers. The perspiration and sublimation fastness is very good to excellent. These are attributed to thermally and chemically stable quinazolinone ring system.

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